

Exothermic	Reaction in which enthalpy decreases $\Delta H = \text{negative}$
Endothermic	Reaction in which enthalpy increases $\Delta H = \text{positive}$
Enthalpy = H	A measure of energy stored in bonds
Entropy = S	A measure of disorder or randomness in a reaction
Free Energy = G	A measure of the spontaneity of a reaction

Definition of an intermediate	A compound formed on the way from starting materials to products along a reaction coordinate. It is at a relative energy minimum.
Definition of a rate determining step (RDS)	The slow step along a reaction coordinate. It is characterized by having the highest activation energy of any of the steps involved in the overall reaction.
Definition of a unimolecular reaction step	A reaction step in which only a single molecule is required to get to the transition state
Definition of a bimolecular reaction step	A reaction step in which two molecules are required to get to the transition state
Definition of a transition state	A structure of high energy along a reaction coordinate that is characterized by partially made and/or broken bonds

Exergonic	Reaction in which free energy decreases $\Delta G = \text{negative}; \text{spontaneous}$
Endergonic	Reaction in which free energy increases $\Delta G = \text{positive}; \text{non spontaneous}$